



# A metal-free visible light active photo-electro-Fenton-like cell for organic pollutants degradation

Dongting Yue<sup>a</sup>, Xufang Qian<sup>a</sup>, Miao Kan<sup>a</sup>, Mengyuan Fang<sup>a</sup>, Jinping Jia<sup>a</sup>, Xudong Yang<sup>b</sup>, Yixin Zhao<sup>a,\*</sup>

<sup>a</sup> School of Environmental Science and Engineering, Shanghai Jiao Tong University, 800 Dongchuan Road, Shanghai 200240, China

<sup>b</sup> State Key Laboratory of Metal Matrix Composites, Shanghai Jiao Tong University, 800 Dong Chuan Road, Minhang District, Shanghai 200240, China

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## ABSTRACT

To avoid the potential secondary metal pollution leached from Fenton-like catalysts and use less or no external chemicals supply, we develop an all-carbon based visible light active photo-electro-Fenton-like (PEF) cell to in-situ generate highly active  $\cdot\text{OH}$  radical for environmental remediation. In this cell, the mesoporous carbon coated graphite felt (MesoC/GF) cathode could effectively produce  $\text{H}_2\text{O}_2$  by electrochemical reduction of oxygen. Graphitic carbon nitride (g-C<sub>3</sub>N<sub>4</sub>) with visible light photocatalytic activities acts as an efficient metal-free Fenton-catalyst for  $\text{H}_2\text{O}_2$  activation to produce  $\cdot\text{OH}$  radical under visible light irradiation. The different cell parameters such as applied voltage, photocatalysts dosage and pH condition have been investigated for phenol removal in aqueous solution. This metal-free visible light active PEF cell shows excellent mineralization efficiency of stubborn phenol with high stability and its performance of phenol removal efficiency is much higher than that of the g-C<sub>3</sub>N<sub>4</sub> only photocatalysis cell and MesoC/GF only electrolysis cell. Moreover, our PEF cell presents comparable or even better Fenton-catalytic activities than the similar electro-Fenton cell using MesoC/GF and homogeneous  $\text{Fe}^{2+}$  ion. These results demonstrate our all-carbon based PEF cell without external chemicals supply is promising for environmental remediation.

## 1. Introduction

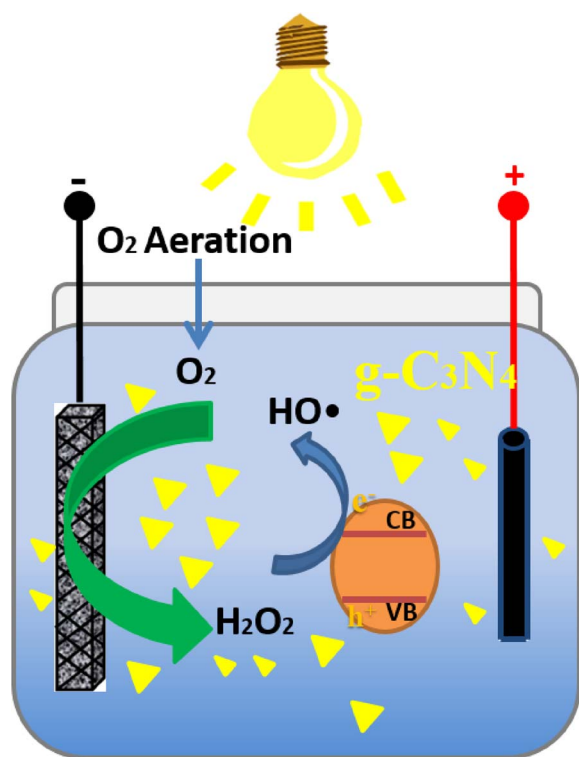
Hydroxyl radical ( $\cdot\text{OH}$ ), as one of the most important reactive oxygen species, has been utilized in various types of environmental remediation including natural waters, the atmosphere, and biological systems [1–5]. With particular attention to its reactivity with various water pollutants, different techniques utilizing hydroxyl radical such as photocatalysis and Fenton reaction have become more and more popular for wastewater treatment [1,6,7]. Recently, advanced oxidation processes based on the generation of  $\cdot\text{OH}$  have been regarded as one of the most effective techniques for the mineralization of persist organic contaminants [6–8]. The classical approach for  $\cdot\text{OH}$  generation is the regular Fenton reaction using  $\text{H}_2\text{O}_2$  and  $\text{Fe}^{2+}$  ion under acidic condition, which is highly active and simple [9]. However, there are two main challenges or problems for homogeneous Fenton reaction: the significant iron residue related secondary pollution and the transportation and storage of  $\text{H}_2\text{O}_2$  with increased safety regulation [9–12]. To overcome these challenges and problems, it is desired to develop some metal-free methods for in-situ generation of  $\cdot\text{OH}$  without external chemicals supply. One promising strategy is to in-situ generate  $\text{H}_2\text{O}_2$

through electrocatalysis process and then activate the  $\text{H}_2\text{O}_2$  into  $\cdot\text{OH}$  via a Fenton-like catalysts with less or no metal ions leakage [13–16].

Among different cathode candidates for electrochemical  $\text{H}_2\text{O}_2$  generation, the low cost carbon materials with advantages of non-toxicity, excellent conductivity and good stability are promising for high yield of  $\text{H}_2\text{O}_2$  [17–19]. The graphite materials such as graphite felt with low cost and high mechanic strength could be a cathode for  $\text{H}_2\text{O}_2$  generation without any possible metals leaching but exhibited relatively low activities. The mesoporous carbon could improve the efficiency of electrochemical  $\text{H}_2\text{O}_2$  generation by facilitating the diffusion and storage of oxygen because of its mesoporous structure. However, it is relatively difficult to process into cathodes [20]. In order to achieve the best balance between performance and cost, here we developed an ordered mesoporous carbon (MesoC) coated commercial graphite felt (GF) cathode (noted as MesoC/GF) for  $\text{H}_2\text{O}_2$  generation via electrochemical reduction of oxygen [20,21]. The catalytic activation of  $\text{H}_2\text{O}_2$  into  $\cdot\text{OH}$  radical is key for Fenton-like reaction, most reported Fenton-like catalysts are based on metal elements to achieve the redox couple for  $\text{H}_2\text{O}_2$  activation. However, it is so difficult to overcome the metal leaching problem completely when using these metal-based Fenton-like

\* Corresponding author.

E-mail address: [yixin.zhao@sjtu.edu.cn](mailto:yixin.zhao@sjtu.edu.cn) (Y. Zhao).



**Scheme 1.** Schematic illustration of visible light active Photo-Electro-Fenton-like cell based on metal-free carbon materials.

catalysts. Here we adapted a well-known metal-free photocatalyst of  $g\text{-C}_3\text{N}_4$  as the visible light active Fenton-like catalyst for activation of  $\text{H}_2\text{O}_2$  into  $\cdot\text{OH}$  [22–32].

Based on these all-carbon materials, we constructed a novel visible light active PEF cell as illustrated in Scheme 1. The highly active MesoC/GF cathode could generate  $\text{H}_2\text{O}_2$  by electrochemical reduction of oxygen, the visible light active  $g\text{-C}_3\text{N}_4$  then activates  $\text{H}_2\text{O}_2$  into  $\cdot\text{OH}$ . In this PEF cell, issues related to both external chemicals supply and metals leaching can be solved. This configuration not only inherits the advantages of traditional electro-Fenton or photo-Fenton generating the  $\cdot\text{OH}$  for the mineralization of most organic pollutant but also has distinct advantages of avoid external chemicals supply and eliminating the metals leaching related environmental risk.

## 2. Experimental

### 2.1. Chemicals and reagents

Graphite Felt (GF) was purchased from Jinglongtetan Technology Co. Ltd. Phenol ( $\geq 99.0$  wt %), ethanol, sodium sulfate, iron(III) nitrate nonahydrate, urea, and methanol were purchased from Sinopharm Chemical Reagent Co., Ltd. Sodium hydroxide, hydrochloric acid and hydrogen peroxide (30 wt%) were obtained from Shanghai Lingfeng Chemical Reagent Co., Ltd. Poly(propylene oxide)-blockpoly(ethylene oxide)-block-poly(propylene oxide) triblock copolymer Pluronic F127 ( $\text{PEO}_{106}\text{PPO}_{70}\text{PEO}_{106}$ , Mw = 12 600), 5,5-dimethylpyrroline-1-oxide (DMPO) was obtained from Tokyo Chemical Industry Co., Ltd. All the aqueous solutions were prepared by using distilled and deionized water.

### 2.2. Preparation of MesoC/GF cathode and $g\text{-C}_3\text{N}_4$

MesoC/GF samples were fabricated by a solvent evaporation-induced self-assembly method with copolymer F127 as a template in an ethanol solution. The resol precursor was prepared according to

previous reported methods [33,34]. In a typical preparation, 1.0 g of F127 was first dissolved in 8.0 g ethanol followed by adding 2.0 g 20 wt % resols ethanol solution with stirring for 10 min until forming a homogeneous solution. The commercial GF ( $3.0\text{ cm} \times 3.5\text{ cm}$ ) was then soaked in the above solution until saturated. The as-made MesoC/GF samples were then annealed at  $350^\circ\text{C}$  for 3 h and  $800^\circ\text{C}$  for 2 h under  $\text{N}_2$  atmosphere.

$g\text{-C}_3\text{N}_4$  was prepared by a simple thermal polycondensation of urea [25,35]. 10 g urea placed in a crucible with a cover was calcinated at  $550^\circ\text{C}$  for 2 h with  $4^\circ\text{C min}^{-1}$  heating rate. Once cooling down, the yellowish  $g\text{-C}_3\text{N}_4$  was the grinded into powder.

### 2.3. Characterization

The crystal phase of the samples were using a Shimadzu XRD-6100 diffractometer with Cu K $\alpha$ . The morphological of these samples were characterized by TEM (JEM-2100F) and FESEM (JSM 7800F). Diffuse reflectance spectra (DRS) were measured on a UV-vis spectrophotometer (Shimadzu UV-2450, Japan). Electron paramagnetic resonance (EPR) spectra was taken on a Bruker EMX-8/2.7C using DMPO as spin trapping agent at 3350 G. The electrochemical test was conducted on CHI 660D electrochemical workstation using the standard Ag/AgCl reference electrode and the graphite counter electrode in 0.05 M  $\text{Na}_2\text{SO}_4$  solution with 50 mL/min  $\text{O}_2$  purging.

### 2.4. Photo-electro-Fenton-like cell

The  $3.0\text{ cm} \times 3.5\text{ cm}$  size MesoC/GF acted as the cathode and graphite electrode as the anode in a two-electrode configuration with 50 mL/min  $\text{O}_2$  purging under different applied voltages. A 100 W LED lamp (Ceaulight, CEL-LED-100) with double 420 nm cutoff filter was used as light source and the irradiation intensity was  $100\text{ mW cm}^{-2}$ . The Photo-Electro-Fenton-like cell contained 70 mL 0.05 M  $\text{Na}_2\text{SO}_4$  aqueous solution with 50 mg/L phenol as model pollutant and different amount of  $g\text{-C}_3\text{N}_4$ . The phenol removal samples were taken at given reaction time using  $0.45\text{ }\mu\text{m}$  membrane filter, and the degradation products and phenol were analyzed on a high-performance liquid chromatography (HPLC, SPD-16, Shimadzu Co., Japan) with a UV/vis detector (270 nm for phenol). Chemical oxygen demand (COD) of the final solution was measured following by China National Standard Method GB11914-89. For the homogeneous system of MesoC/GF electrolysis cell with  $\text{Fe}^{2+}$ , excess methanol was added into samples that were removed from the cell at specific time intervals for quenching of the reaction. The  $\text{H}_2\text{O}_2$  concentration was colorimetrically analyzed at 410 nm on the UV spectrophotometer (Cary 60 UV-Vis) using  $\text{TiOSO}_4/\text{H}_2\text{SO}_4$  complexation reagent [11,36].

### 2.5. Results and discussions

The photographs of the commercial GF cathode and the mesoporous carbon modified MesoC/GF cathode are shown in Fig. 1a. The MesoC/GF with mesoporous carbon coating looks darker than the bare GF. The SEM image shows that the commercial GF exhibits fiber and belt morphology with a smooth surface, while the surface of MesoC/GF becomes uneven and rough after modified with mesoporous carbon (Fig. 1b and c). These mesoporous carbon materials penetrated within the GF frameworks. TEM image of the MesoC/GF in Fig. 1d presents the strip like mesopore channels, confirming the ordered hexagonal mesostructure of mesoporous carbon in MesoC/GF cathode [33,37]. These observations are consistent with the fact that the MesoC/GF has a larger BET surface area than the GF (Fig. S1). The electro-catalytic performance of GF and MesoC/GF electrodes for oxygen reduction is presented in Fig. 1e and f. The GF electrode shows two reduction peaks at  $-0.58\text{ V}$  and  $-1.2\text{ V}$  vs Ag/AgCl. The reduction peaks of MesoC/GF electrode are negatively shifted to  $-0.42\text{ V}$  and  $-0.86\text{ V}$  vs Ag/AgCl, respectively. These results demonstrate that MesoC/GF cathode

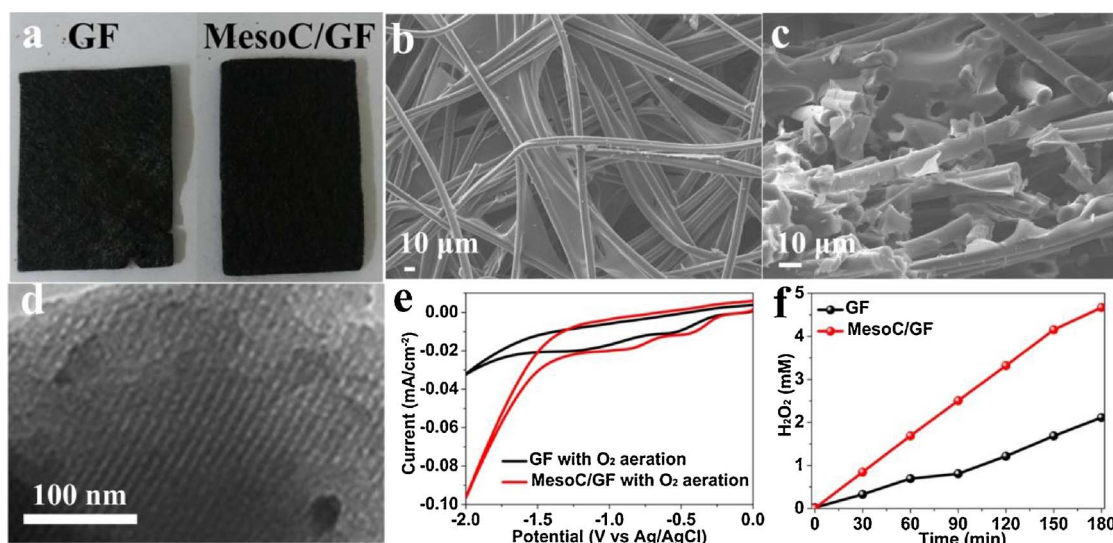


Fig. 1. Photographs (a) and SEM images of GF (b) and MesoC/GF (c), TEM image of MesoC/GF cathode (d), Cyclic voltammograms curve of GF and MesoC/GF cathodes at pH = 3.0 (e),  $\text{H}_2\text{O}_2$  generation from GF and MesoC/GF cathodes at  $-0.8\text{ V}$  vs Ag/AgCl and pH = 3.0 (f).

presents the higher electro-catalytic activities for oxygen reduction than GF cathode. Moreover, the rate and amount of  $\text{H}_2\text{O}_2$  generation by MesoC/GF cathode at  $-0.8\text{ V}$  vs Ag/AgCl is much higher than GF cathode (Fig. 1f), in which  $\text{H}_2\text{O}_2$  concentration increases from 2.1 mM to 4.9 mM after 180 min. The modification of commercial GF with MesoC significantly enhanced oxygen reduction activities to generate  $\text{H}_2\text{O}_2$ . It is believed that the introduction of MesoC on GF could facilitate both the electron transfer and the diffusion and storage of oxygen, which would efficiently improve the reduction efficiency of oxygen to produce  $\text{H}_2\text{O}_2$  on the cathode [38,39]. Consequently, the MesoC/GF cathode is promising for electrochemical generation of  $\text{H}_2\text{O}_2$ .

The metal-free  $\text{g-C}_3\text{N}_4$  was facilely synthesized and simply grinded as micron size particles, which was easier to separate and recycle than small nanoparticles. XRD patterns of the  $\text{g-C}_3\text{N}_4$  catalyst presents two distinct diffraction peaks at  $13.2^\circ$  and  $27.5^\circ$ , which is well indexed to  $\text{g-C}_3\text{N}_4$  (JCPDS Card No.87–1526) (Fig. 2a). The  $\text{g-C}_3\text{N}_4$  shows an absorption onset edge at  $\sim 580\text{ nm}$  in Fig. 2b and the corresponding band gap of the  $\text{g-C}_3\text{N}_4$  is about  $2.7\text{ eV}$  [40–42]. SEM image in Fig. 2c indicates that  $\text{g-C}_3\text{N}_4$  possesses plate-like structures, which is beneficial for its separation and recycle in real applications. According to previous reports [26,27], the calculated conduction band position of  $\text{g-C}_3\text{N}_4$  is  $\sim 0.88\text{ V}$  vs. NHE, while the valence band is  $\sim 2.14\text{ V}$  vs. NHE. Under visible light irradiation, the  $\text{g-C}_3\text{N}_4$  could generate photoelectrons with potential of  $\sim 0.88\text{ V}$  vs. NHE, which have sufficient force to activate  $\text{H}_2\text{O}_2$  to form  $\cdot\text{OH}$  and  $\text{OH}^-$ . The energy level diagram for the Fenton activation and the electronic structure of  $\text{g-C}_3\text{N}_4$  is illustrated in Fig. S2. The products of  $\cdot\text{OH}$  and holes are the strong oxidants and could be used for removing organic contaminants.

To simplify the device configuration and reduce instrument cost, we

constructed a two electrodes configuration using a DC power supply to drive electrochemical  $\text{H}_2\text{O}_2$  generation. The  $\text{g-C}_3\text{N}_4$  was then suspended in the solution as the photocatalytic Fenton-like catalyst for  $\text{H}_2\text{O}_2$  activation, as shown in Scheme 1. Under the  $1.0\text{ V}$  applied voltage, there is negligible  $\text{H}_2\text{O}_2$  found in the solution (Fig. 3a) because the applied voltage is not high enough to drive the electrochemical oxygen reduction. Once the applied voltages are increased to  $3\text{ V}$  and  $5\text{ V}$ , a large amount of  $\text{H}_2\text{O}_2$  are generated in the solutions. The stable  $\text{H}_2\text{O}_2$  concentration could be up to  $0.9\text{ mM}$  and  $1.6\text{ mM}$ , respectively (Fig. 3a). Here, the  $\text{H}_2\text{O}_2$  concentration in the PEF cell is lower than the electrolysis only cell because the generated  $\text{H}_2\text{O}_2$  is consumed by the Fenton-like reaction. The phenol removal efficiency in PEF cell was also investigated with varied parameters. As shown in Fig. 3b, the phenol degradation efficiency by PEF cell was remarkably enhanced with the increase of the applied voltage from  $1\text{ V}$  (6.3%) to  $3\text{ V}$  (91.1%) when the amount of  $\text{g-C}_3\text{N}_4$  and phenol was fixed. However, the phenol removal efficiency decreased to 74.2% as the applied voltage was increased to  $5\text{ V}$ . This could be due to the variation of  $\text{H}_2\text{O}_2$  generated at the MesoC/GF cathode at different applied voltages (Fig. 3a). Although the higher amount of  $\text{H}_2\text{O}_2$  can be generated by MesoC/GF cathode at applied voltage of  $5.0\text{ V}$ , the degradation efficiency for phenol at  $5.0\text{ V}$  applied voltage decreases and is lower than applied voltage of  $3.0\text{ V}$ . This observation could be ascribed to the mechanism that excessive  $\text{H}_2\text{O}_2$  could act as a self-scavenger for  $\cdot\text{OH}$ , following Eqs. (1) and (2), which leads to the greatly decrease of  $\cdot\text{OH}$  generation [43].



Here, we fixed the applied voltage as  $3\text{ V}$  for electrochemical  $\text{H}_2\text{O}_2$

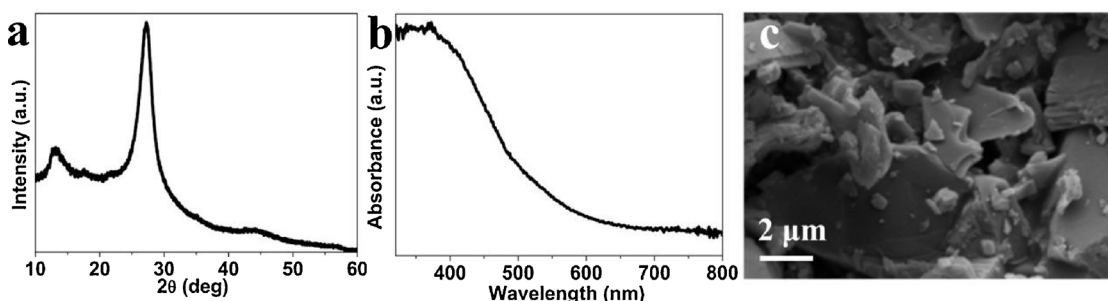


Fig. 2. XRD pattern (a), UV-Vis diffuse reflectance spectrum (b), and SEM images of  $\text{g-C}_3\text{N}_4$  (c).



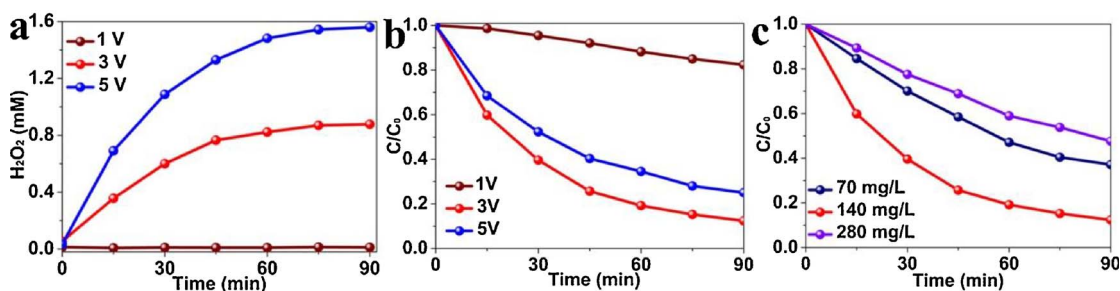


Fig. 3. H<sub>2</sub>O<sub>2</sub> generation (a) and phenol degradation (b) in PEF cell at different applied voltage using 140 mg/L g-C<sub>3</sub>N<sub>4</sub> at pH = 4.0. The effect of different g-C<sub>3</sub>N<sub>4</sub> dosage on phenol degradation with applied voltage 3 V and pH = 4.0 (c).

generation in the following paragraph, the effect of g-C<sub>3</sub>N<sub>4</sub> dosage on phenol removal efficiency was also studied. Fig. 3c reveals that the phenol removal efficiency first increases when the dosage of g-C<sub>3</sub>N<sub>4</sub> increases from 70 mg/L to 140 mg/L, but then decreases once the amount of g-C<sub>3</sub>N<sub>4</sub> increases to 280 mg/L. The decrease of phenol removal efficiency with increase of g-C<sub>3</sub>N<sub>4</sub> could be due to the light scattering and light blocking effects by too much g-C<sub>3</sub>N<sub>4</sub>. Additionally, since some nitrogen-containing by-product may be generated from g-C<sub>3</sub>N<sub>4</sub> especially under acidic conditions (pH = 4), leading to the exceeding of total nitrogen concentrations. Total nitrogen (TN) concentrations with different dosage of g-C<sub>3</sub>N<sub>4</sub> in phenol solution at pH 3 were investigated in Table S1. These results demonstrated that TN (30.2 mg/L) of 280 mg/L g-C<sub>3</sub>N<sub>4</sub> dosage exceeded the standard of TN emission concentration (20.0 mg/L), while TN (14.4 mg/L) of 140 mg/L g-C<sub>3</sub>N<sub>4</sub> dosage meet the wastewater regulation. As a result, 140 mg/L of g-C<sub>3</sub>N<sub>4</sub> was selected as an optimal photocatalyst dosage for  $\cdot$ OH activation or phenol removal in the following studies. Therefore, the applied voltage of 3.0 V and 140 mg/L g-C<sub>3</sub>N<sub>4</sub> was used for the following investigations. The above results have shown that the PEF cell integrating MesoC/GF cathode for in-situ H<sub>2</sub>O<sub>2</sub> generation and g-C<sub>3</sub>N<sub>4</sub> as visible light active Fenton-like catalyst could effectively degrade phenol using electricity and visible light without external chemicals supply.

To reveal whether the Fenton-like reaction or other mechanism contribute to the phenol degradation, we compare our PEF cell with the g-C<sub>3</sub>N<sub>4</sub> only photocatalysis cell and the MesoC/GF only electrolysis cell for degradation of phenol. The MesoC/GF only electrolysis cell with large amounts of H<sub>2</sub>O<sub>2</sub> generation presents a lower phenol removal efficiency of 34.8% after 180 min (Fig. 4a). Furthermore, the HPLC results in Fig. 4b revealed that the observed phenol removal was not the complete degradation but only transformation of phenol into some other organic intermediates (Fig. 4b). This result suggested that the H<sub>2</sub>O<sub>2</sub> cannot completely degrade phenol but just oxidize it to form some intermediate due to the difficulty of breaking down the benzene ring. In contrast, the phenol removal by g-C<sub>3</sub>N<sub>4</sub> only photocatalysis system is just 20.1% but the observed phenol removal is mainly mineralized, which is degraded by the active  $\cdot$ O<sub>2</sub><sup>-</sup> and h<sup>+</sup> via the photocatalysis process (Fig. 4b). After coupling the photocatalysis of g-C<sub>3</sub>N<sub>4</sub> with MesoC/GF electrolysis cell as PEF cell, the same amount of phenol was rapidly degraded with a removal efficiency of 99.1%. Although most phenol was mineralized into CO<sub>2</sub> and H<sub>2</sub>O, some by-product organic compounds including oxalic acid and formic acid were also generated during the degradation process of phenol (Fig. S3). Consequently, a plausible pathways of phenol degradation in the visible active PEF cell was proposed in Fig. S4. Additionally, some other persist organic pollutants including tetracycline and 2,4,6-trichlorophenol were also investigated. Results in Fig. S5 demonstrated that the tetracycline removal efficiency was reached to 90% and complete degradation of 2,4,6-trichlorophenol was achieved within 180 min, suggesting the high activities of PEF cell for persist organic pollutants degradation.

The concentrations of H<sub>2</sub>O<sub>2</sub> generated in these different systems were analyzed to explore the different performances of these systems for phenol degradation (Fig. 4c). There was almost no H<sub>2</sub>O<sub>2</sub> detected in

the g-C<sub>3</sub>N<sub>4</sub> only photocatalysis, although some previous reports suggested that g-C<sub>3</sub>N<sub>4</sub> could produce small amount of H<sub>2</sub>O<sub>2</sub> via photocatalysis process. The accumulated H<sub>2</sub>O<sub>2</sub> (3.5 mM) in MesoC/GF only electrolysis cell increased rapidly in the first 60 min then grew slowly. The H<sub>2</sub>O<sub>2</sub> generated in MesoC/GF only electrolysis cell cannot spontaneously become highly active  $\cdot$ OH, which leads to the lower phenol degradation efficiency with larger amount of H<sub>2</sub>O<sub>2</sub>. In contrast, the relatively lower concentration of H<sub>2</sub>O<sub>2</sub> (0.85 mM) after 60 min in PEF cell was attributed to that most in-situ generated H<sub>2</sub>O<sub>2</sub> were effectively converted into  $\cdot$ OH via photogenerated electrons catalytic process. These results are consistent with the above mentioned observation that the phenol removal efficiency via PEF cell is much higher than g-C<sub>3</sub>N<sub>4</sub> only photocatalysis cell and MesoC/GF only electrocatalysis cell. Based on above mentioned results, the role of g-C<sub>3</sub>N<sub>4</sub> in our PEF cell could not only form the active oxygen species like  $\cdot$ O<sub>2</sub><sup>-</sup> and highly active h<sup>+</sup> via the regular photocatalysis process but also produce  $\cdot$ OH by activating in-situ generated H<sub>2</sub>O<sub>2</sub>, all of which contribute to the high phenol removal efficiency and high mineralization efficiency.

Since the pH values can significantly impact the Fenton reaction, it is necessary to determine the effect of initial pH on the phenol degradation efficiencies (Fig. 5). A notable pH effect was observed in the PEF cell for degradation of phenol. At pH = 3, 99.1% of phenol was degraded within 180 min, while the phenol removal efficiency was 80.2% and 71.3% at pH = 5.0 and 7.0 respectively (Fig. 5a). Simultaneously, the pH variation has no obvious effect on electrochemical generation of H<sub>2</sub>O<sub>2</sub> (Fig. 5b). It seems that phenol removal efficiency under different pH values is mainly determined by the activation of  $\cdot$ OH (Eq. (3)) rather than the generation of H<sub>2</sub>O<sub>2</sub> (Eq. (4)), although both of which can be affected by pH value. The reason for pH independent electrochemical generation of H<sub>2</sub>O<sub>2</sub> in the PEF cell could be that the O<sub>2</sub> diffusion is the limitation reaction step for the O<sub>2</sub> reduction.



The above conclusion was also further confirmed by some control experiments, in which we used the g-C<sub>3</sub>N<sub>4</sub> as a photocatalytic Fenton-like catalyst in 0.85 mM H<sub>2</sub>O<sub>2</sub> solution containing 50mg/L phenol. Fig. S6a illustrates that the lower solution pH leads to the enhancement of phenol removal efficiency increases from 18.3% to 40.1%, respectively, within 90 min. Correspondingly, H<sub>2</sub>O<sub>2</sub> concentration generally decreases with reaction time, and the amount of H<sub>2</sub>O<sub>2</sub> at pH = 3, 5 and 7 is reduced from 0.85 mM to 0.69 mM, 0.76 mM and 0.83 mM, respectively (Fig. S6b). The formation of  $\cdot$ OH catalyzed via photoelectron (Eq. (3)) is facilitated under acidic condition, which is consistent with the highest phenol degradation efficiency at pH = 3. Here the phenol removal efficiency using g-C<sub>3</sub>N<sub>4</sub> as photocatalysis with added H<sub>2</sub>O<sub>2</sub> is lower than our system incorporated with MesoC/GF for in-situ H<sub>2</sub>O<sub>2</sub> generation, which could be ascribed to the factor that the H<sub>2</sub>O<sub>2</sub> decreases with reaction and the absence of oxygen bubbling could also reduce the activities of g-C<sub>3</sub>N<sub>4</sub>.

As shown in Fig. 6a, PEF cell presents the excellent degradation of

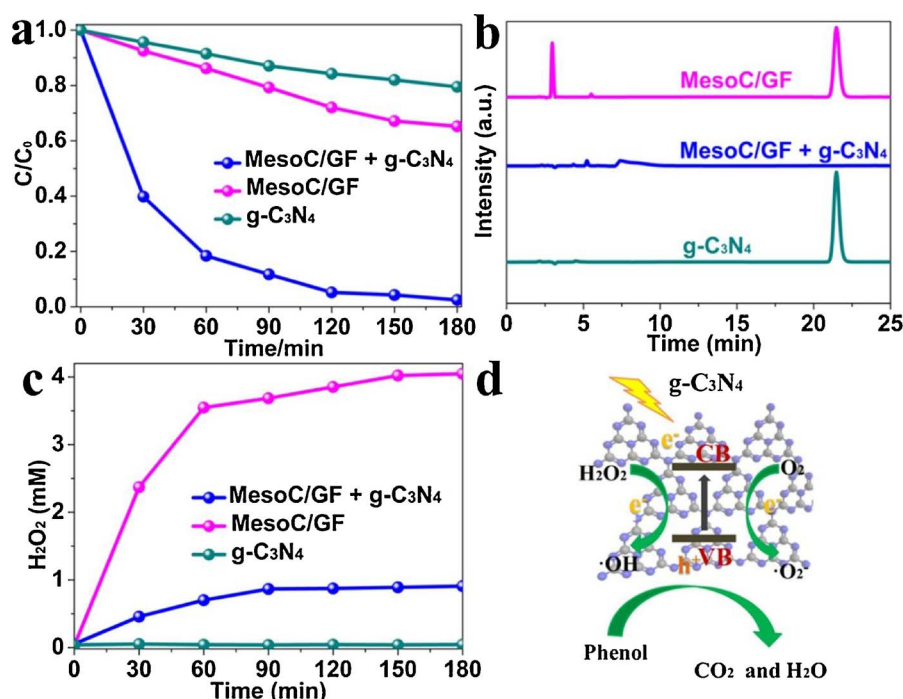


Fig. 4. The phenol degradation (a), time-dependent evolution of the HPLC diagrams for phenol degradation after 180 min (b), and the  $\text{H}_2\text{O}_2$  yield (c) in the different systems: PEF cell based on MesoC/GF + g-C<sub>3</sub>N<sub>4</sub>, electrolysis cell only based on MesoC/GF, and g-C<sub>3</sub>N<sub>4</sub> only photocatalysis cell. Schematic illustration of phenol degradation mechanism in PEF cell based on MesoC/GF + g-C<sub>3</sub>N<sub>4</sub> (d).

phenol over a wide pH range of 3.0–7.0 and the best performance is achieved at pH = 3. The most exciting point is that it is comparable to the traditional Electro-Fenton cell using 2 mg/L  $\text{Fe}^{2+}$  as  $\text{H}_2\text{O}_2$  activator when the same MesoC/GF cathode is used as  $\text{H}_2\text{O}_2$  generator with 3 V applied voltage at pH = 3. The phenol removal efficiency in homogeneous Electro-Fenton cell based on the MesoC/GF electrolysis cell using 2 mg/L  $\text{Fe}^{2+}$  at neutral pH = 7 decreases to only 30.2%, while our PEF cell still exhibits 72.0% phenol removal. This is a significant enhancement on the phenol removal efficiency under neutral condition by using our PEF cell. The DMPO spin-trapping EPR spectra of our PEF cell is listed in Fig. 6b, which displaces four typical DMPO-OH adducts with intensity of 1:2:2:1 as  $\cdot\text{OH}$ . The EPR result confirms the formation of  $\cdot\text{OH}$  via Fenton-like photocatalyst of g-C<sub>3</sub>N<sub>4</sub> under visible light irradiation. The relative EPR signals intensity decrease at lower pH in both PEF cell and the MesoC/GF electrolysis cell with  $\text{Fe}^{2+}$ . While PEF cell shows much higher intensity of DMPO – OH adduct signal than the case of the MesoC/GF electrolysis cell in the presence of  $\text{Fe}^{2+}$  at pH = 7. The EPR results are in consistent with the trend of removal efficiency at different pH values, suggesting that the efficient generation of  $\cdot\text{OH}$  in PEF cell for the phenol degradation. From the perspective of actual applications, the PEF cell's stability is an important issue that must be considered other than high activities. With the cycle tests, our PEF cell shows almost the same catalytic activity for the phenol decomposition. The COD removal efficiency for all the cycle tests are

more than 85% as shown in Fig. 6c and d. To further check the structural stability of g-C<sub>3</sub>N<sub>4</sub> catalyst, the crystal structure and morphology of g-C<sub>3</sub>N<sub>4</sub> after reaction and annealing treatment were investigated by XRD and SEM. The phase and bulk structure of g-C<sub>3</sub>N<sub>4</sub> were maintained similar to those of pristine sample, even the cycle tests (Figure S7 and S8). All these results indicate the high mineralization ability and stability of the PEF cell.

### 3. Conclusions

In summary, a novel metal-free PEF cell integrating a functional carbon based cathode for  $\text{H}_2\text{O}_2$  generation with a carbon based photocatalytic Fenton-like catalyst of g-C<sub>3</sub>N<sub>4</sub> would not only provide the oxidizing agents including  $\text{h}^+$  and  $\cdot\text{O}_2^-$  through the photocatalysis but also in-situ generate  $\cdot\text{OH}$  with high activities for environmental remediation. The secondary pollution related to metal residue is completely eliminated in this system. Our work will push forward the application of carbon based cathode materials and visible light driven catalysts for solar energy assisted water treatment and environmental remediation.

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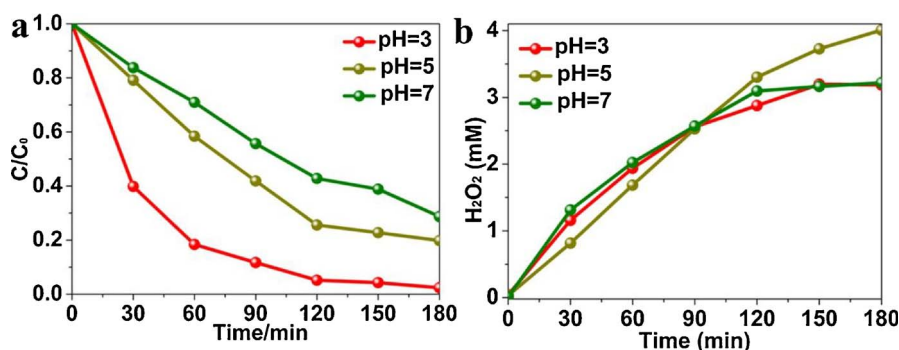


Fig. 5. Phenol degradation by our PEF cell at the different pH with applied voltage 3 V and 140 mg/L g-C<sub>3</sub>N<sub>4</sub> (a), Effect of pH on  $\text{H}_2\text{O}_2$  generation in the absence of g-C<sub>3</sub>N<sub>4</sub> (b).

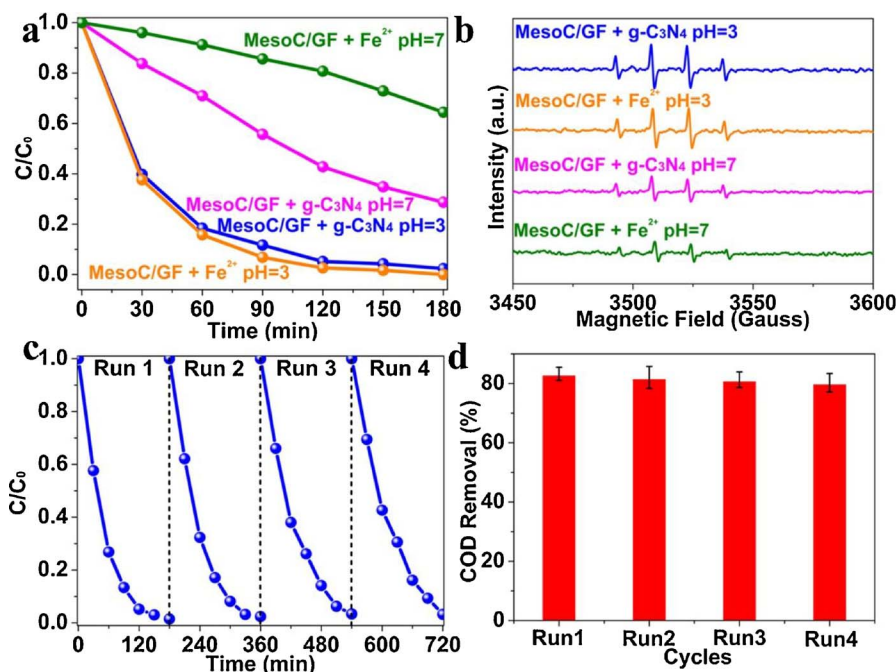


Fig. 6. Effects of pH value on the phenol degradation (a) and DMPO spin-trapping EPR spectra of different pH value on activation (b) via the PEF cell and the MesoC/GF electrolysis cell in the presence of 2 mg/L  $\text{Fe}^{2+}$ . Recycling test of phenol degradation (c) and COD removal efficiency (d) in the PEF cell at pH = 3.0.

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## Appendix A. Supplementary data

Supplementary data associated with this article can be found, in the online version, at <https://doi.org/10.1016/j.apcatb.2018.02.033>.

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